

**TOPIC**

Chemistry – Section III – Question 5

**QUESTION**

The boiling point elevation constant for water is  $K_{bp} = 0.52^{\circ}\text{C/molal}$ . If the normal boiling point of water is  $100^{\circ}\text{C}$ , the normal boiling point (in  $^{\circ}\text{C}$ ) of water containing 1 % by weight NaCl (assume NaCl is completely dissociated) most nearly is

- (A) 100.09
- (B) 100.18
- (C) 105.20
- (D) 110.40

**HINT**

Molality of the ions will be twice the molality of the salt if it is completely dissociated.

**CONTRIBUTOR**

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